

Constituent	AS (µg/L)	CS (µg/L)
Arsenic (trivalent, dissolved)	$360 \times 1.0^* = 360$	$190 \times 1.0^* = 190$
Boron (total)	40,100	7,600

Cadmium (dissolved)	$e^{A+B\ln(H)} \times \left\{ 1.138672 - \left[ \frac{1.138672 - 1.101672}{[(\ln(H))(0.041838)]} \right] \right\}^*$	$e^{A+B\ln(H)} \times \left\{ 1.101672 - \left[ \frac{1.101672 - 1.074672}{[(\ln(H))(0.041838)]} \right] \right\}^*$
	where $A = -2.918$ and $B = 1.128$	where $A = -3.490$ and $B = 0.7852$
Chromium (hexavalent, total)	16	11
Chromium (trivalent, dissolved)	$e^{A+B\ln(H)} \times 0.316^*$	$e^{A+B\ln(H)} \times 0.860^*$
	where $A = 3.688$ and $B = 0.8190$	where $A = 1.561$ and $B = 0.8190$
Copper (dissolved)	$e^{A+B\ln(H)} \times 0.960^*$	$e^{A+B\ln(H)} \times 0.960^*$
	where $A = -1.464$ and $B = 0.9422$	where $A = -1.465$ and $B = 0.8545$
Cyanide**	22	5.2
Fluoride (total)	$e^{A+B\ln(H)}$	$e^{A+B\ln(H)}$ , but <b>must</b> <b>shall</b> not exceed 4.0 mg/L
	where $A = 6.7319$ and $B = 0.5394$	where $A = 6.0445$ and $B = 0.5394$
Lead (dissolved)	$e^{A+B\ln(H)} \times \left\{ 1.46203 - \left[ \frac{1.46203 - 1.42503}{[(\ln(H))(0.145712)]} \right] \right\}^*$	$e^{A+B\ln(H)} \times \left\{ 1.46203 - \left[ \frac{1.46203 - 1.42503}{[(\ln(H))(0.145712)]} \right] \right\}^*$
	where $A = -1.301$ and $B = 1.273$	where $A = -2.863$ and $B = 1.273$
Manganese (dissolved)	$e^{A+B\ln(H)} \times 0.9812^*$	$e^{A+B\ln(H)} \times 0.9812^*$

	where $A = 4.9187$ and $B = 0.7467$	where $A = 4.0635$ and $B = 0.7467$
Mercury (dissolved)	$2.6 \times 0.85^* = 2.2$	$1.3 \times 0.85^* = 1.1$
Nickel (dissolved)	$e^{A+B \ln(H)} \times 0.998^*$	$e^{A+B \ln(H)} \times 0.997^*$
	where $A = 0.5173$ and $B = 0.8460$	where $A = -2.286$ and $B = 0.8460$
TRC	19	11
Zinc (dissolved)	$e^{A+B \ln(H)} \times 0.978^*$	$e^{A+B \ln(H)} \times 0.986^*$
	where $A = 0.9035$ and $B = 0.8473$	where $A = -0.4456$ and $B = 0.8473$
Benzene	4200	860
Ethylbenzene	150	14
Toluene	2000	600
Xylene(s)	920	360

where:

$\mu\text{g/L}$  = microgram per liter

$e^x$  = base of natural logarithms raised to the x-power

$\ln(H)$  = natural logarithm of hardness (in mg/L as  $\text{CaCO}_3$ ) **Hardness**

\* = conversion factor multiplier for dissolved metals

\*\* = standard to be evaluated using either of the following USEPA approved methods, incorporated by reference at 35 Ill. Adm. Code 301.106:  
Method OIA-1677, DW: Available Cyanide by

Flow Injection, Ligand Exchange, and  
Amperometry, January 2004, Document Number  
EPA-821-R-04-001 or Cyanide Amenable to  
Chlorination, Standard Methods 4500-CN-G (40  
CFR 136.3)